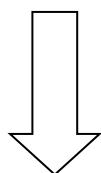


Review Sheet Chapter 19 Acids/Bases	Mrs Sanford	Name:
Properties of Acids-		
Properties of Bases-		
Monoprotic-	Diprotic-	
Arrhenius Acid-		
Arrhenius Base-		
Lewis diagram and structural diagram of hydronium ion-	Amphoteric-	
Bronsted Lowry Acid:	Bronsted Lowry Base:	
Bronsted/Lowry -give a reaction example with A and B labeled-		
Conjugate acid /base pairs-Give an example:		
The stronger the acid,the_____its conjugate base.(when looking at the acid/base table)		
Example of a Lewis acid and a Lewis base:		
Self-Ionization of water:		
What does Kw stand for and give an equation for it with the value of the constant:		
In a basic solution:	In an acidic solution:	
pH=	pOH=	14=
Strong Acid:	Weak Acid:	
How does the value of Ka get affected by strength of the acid?		
Ka=	Kb=	
Calculating a Dissociation constant using an ICE table: There is a calculation of this type below. P 610 good review		
Know how to use the indicator table.		
Five steps for predicting acid/base reactions 1-		
2-		
3-		
4-		
5-		
Draw a sketch and label a pH curve with a base as a titrant and two quantitative reactions.		Draw a sketch and label a pH curve with An acid as a titrant and one quantitative Reaction.



(Label Ph endpoint, equivalence point, point of neutralization on your sketch graphs)

What does a strong dilute acid mean?	What does a strong concentrated acid mean?	
Acid base stoichiometry How do you find volume of titrant?	Why is a balanced equation important?	Write an example of a neutralization reaction involving coefficients other than 1.

Do. P 625 #45 to 50 55 to 57 58 to 61 63, 65 p 626 74 a,c, 75a,c,85,
p 610 22, 23(ICE calculations), Give an example for a 5 step with \rightarrow , $>50\%$, and $<50\%$ (an example for each)

Chem 121-Buffers/Titrations involving endpoints/Salt hydrolysis
Salt hydrolysis:
Salts that produce acidic solutions contain:
Salts that produce basic solutions contain:
Strong Acid + _____ \rightarrow Neutral solution Strong acid + _____ \rightarrow acidic solution
Weak acid + _____ \rightarrow Basic Solution
Buffer:
In a buffer the pH remains constant when:
Buffer capacity:
Write a reaction to show: 1) Acid being added to a solution of hydrogen phosphate
2) Base being added to a solution of hydrogen phosphate
3) Write an equation to show what happens when acid is added to the ethanoic acid-ethanoate ion buffer.
Titration calculations with endpoint info. given: Try the question below
A 25.0 mL solution of diluted rust-removing solution containing phosphoric acid was titrated to the second endpoint using 1.50mol/L sodium hydroxide. The average equivalence point of the sodium hydroxide solution was 17.9mL. what is the concentration of the phosphoric acid in the rust-removing solution?